

# PHOSPHATE PART II

<https://www.jpmmc.com.jo/Default.aspx>

Reference: Shreve's Book Ch. 10, pp. 244-261

# Phosphoric Acid Production

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- ❑ Metallic acid with the chemical formula of ( $\text{H}_3\text{PO}_4$ ), also called Orthophosphoric Acid.
- ❑ Orthophosphoric acid used in fertilizers industry.
- ❑ Diluted phosphoric acid of (28%  $\text{P}_2\text{O}_5$ ) concentration are produced daily, depending on the type.
- ❑ By acidulation with Sulfuric Acid, Phosphate rock is converted to Phosphoric Acid and Gypsum is a by product.
- ❑ The phosphoric acid plant produces the following by – products:
  - 25% hexa fluorosilicic acid ( $\text{H}_2\text{SiF}_6$ )
  - Gypsum, 25-30% free water.

# Phosphoric Acid Production

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## Production

- Phosphate is moved by conveyor belts to Phosphate Crusher, where all particles are having the size of less than 500 micron.
- Powdered phosphate reacts with sulfuric acid in a reactor which produced diluted phosphoric acid and gypsum.



- The mixture is, then pumped to three incubators, in order to enlarge the gypsum crystals.

# Phosphoric Acid

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- The resulting solution is filtered, and the diluted phosphoric acid is sent to the acid storage, to be concentrated later.
- The resulting gases from the reaction, which include multiple fluoride compounds, water vapor, and acids, are washed by absorption towers before they are released to the atmosphere.

# Phosphoric Acid

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- Diluted phosphoric acid is concentrated in heat- exchangers from 28% to 52% in three concentration lines, using vacuum evaporation.
- During this concentration process HF reacts with silica (which is present in the rock in sufficient amounts) producing fluorosilicic acid ( $\text{H}_2\text{SiF}_6$ ) in 22% concentration and water vapor.



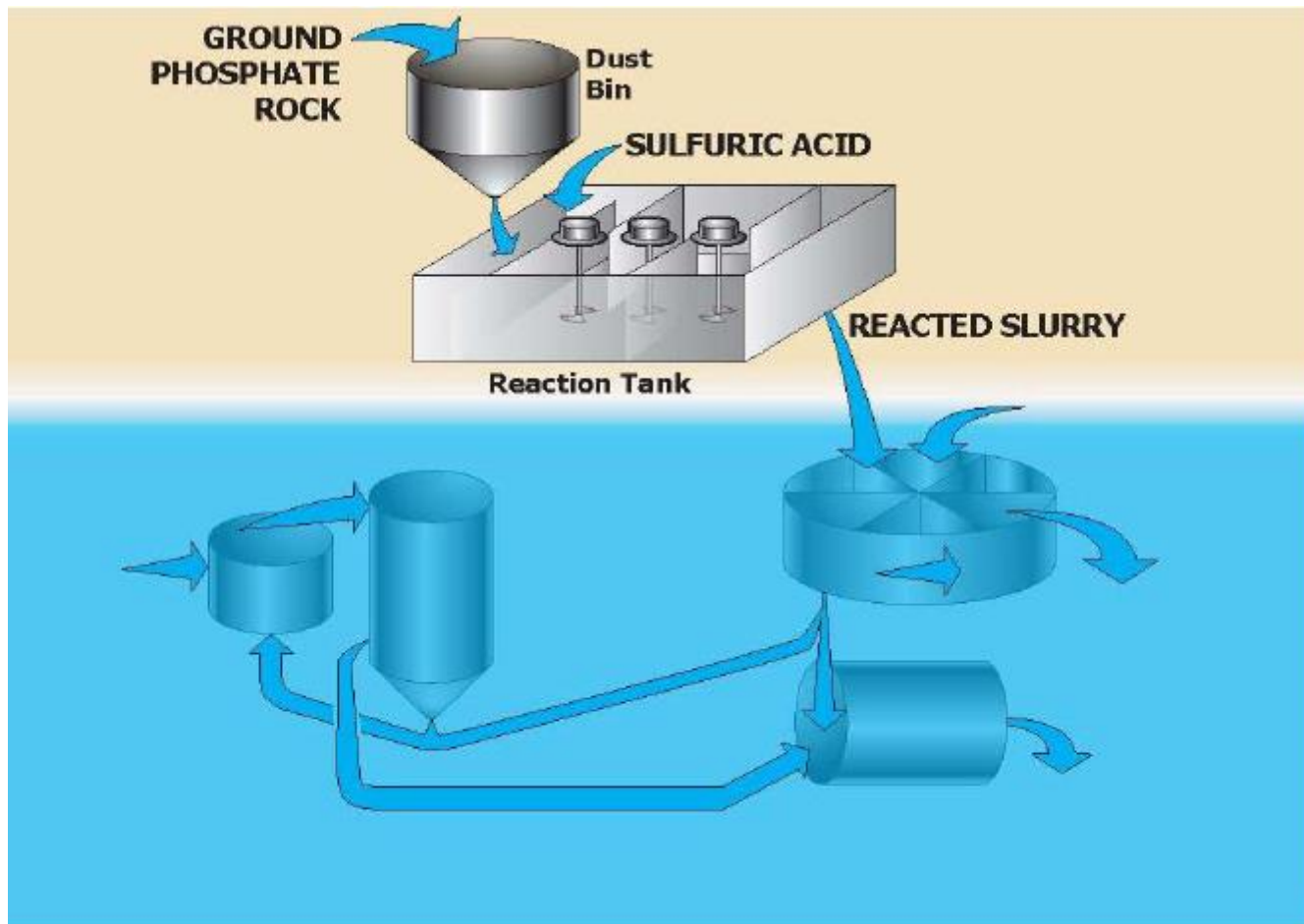
- Purification to part of the acid is applied to have part of the final product for exporting as a final product; a minor part is sent to the local market.

## Usages:

Phosphoric acid is an intermediate product for multiple industries such as fertilizers, animal feed, detergents, and some food industries.

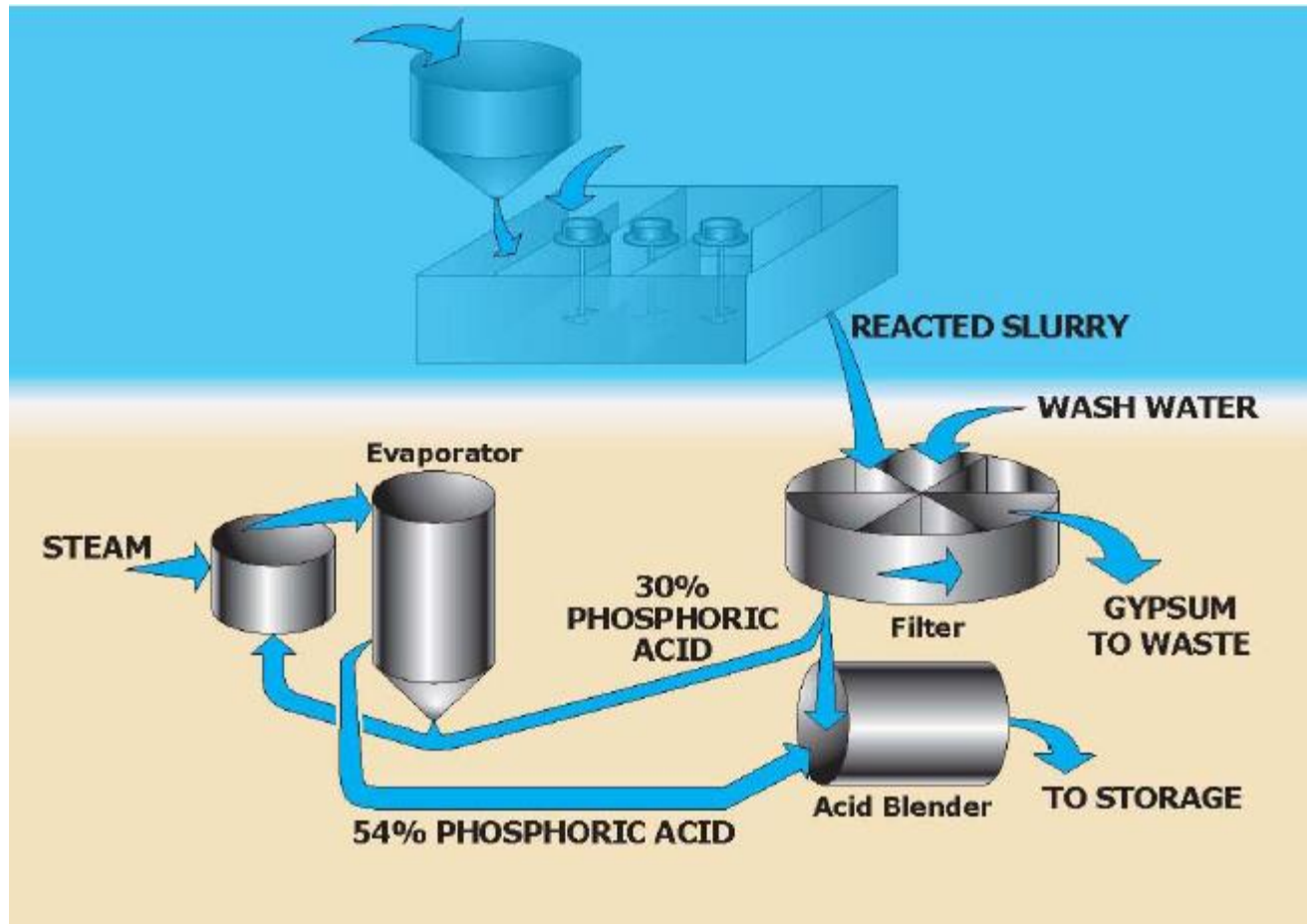
# Phosphoric Acid

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# Phosphoric Acid 📢

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# Environmental Hazards of $\text{H}_3\text{PO}_4$ Production

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## 1. Solid waste

- Phosphogypsum (PG) is a waste byproduct produced by phosphoric acid, about 5 tones of PG are generated for every tone of phosphoric acid ( $\text{H}_3\text{PO}_4$ ) produced.
- Some impurities naturally present in the phosphate rock become concentrated in (PG), including fluoride compounds, heavy metals such as lead and cadmium, radioactive elements and residual acidity.
- The continued accumulation of (PG) has created urgent pressures to find useful applications for this by-product.



# Environmental Hazards of $\text{H}_3\text{PO}_4$ Production

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## 2. Liquid waste

- ❑ The flouroslicic acid shall be supplied to produce aluminum fluoride.
- ❑ Off- grade flouroslicic acid shall be recycled to  $\text{H}_3\text{PO}_4$  plant.
- ❑ Any remaining amount of off- grade flouroslicic acid will be neutralized with limestone to produce calcium silicate.

# Environmental Hazards of $\text{H}_3\text{PO}_4$ Production

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## 3. Emissions

- a. Major emissions from  $\text{H}_3\text{PO}_4$  production include  $\text{SiF}_4$ , HF, fluorine gas.
  - These gases will be collected by ventilation system and the sent to scrubber.
- b. The secondary emission is dust originated from handling, grinding of phosphate rock.
  - Dust collecting equipment will be installed.

# Sulphuric Acid Production

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- It is a strong metallic soluble acid at all concentrations, has a chemical formula  $\text{H}_2\text{SO}_4$ .
- Sulfuric acid is produced in two units with 98.5% concentration, and a designed daily capacity of 2,500 tons for each unit.
- Sulfur used in producing sulfuric acid is imported from Arab neighboring countries, Iraq and Saudi Arabia, and from some foreign countries as Russia and Iran.
- Sulfur is stored in a storage facility of a capacity of 35,000 tons.

# Sulphuric Acid Production

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## Production

- ❑ Sulfur is dragged from the stores by loaders and fed onto conveyor belts, where it gets melted in special pools by medium-pressure vapor, and adding some materials for purification.
- ❑ Lime was used to neutralize sulfur acidity before melting.
- ❑ Liquid sulfur is stored in a heat-insulated storage at a temperature of 135 °C.
- ❑ Liquid sulfur is pumped into a kiln at a temperature of 1000 °C, where it is burnt with the presence of dry atmospheric air to be transformed into sulfur dioxide in the state of gas.
- ❑ As that chemical reaction is an exothermal one, the released temperature is then used in producing high pressure steam (at 46 atm), which conveyed to the Facilities Unit.

# Sulphuric Acid Production

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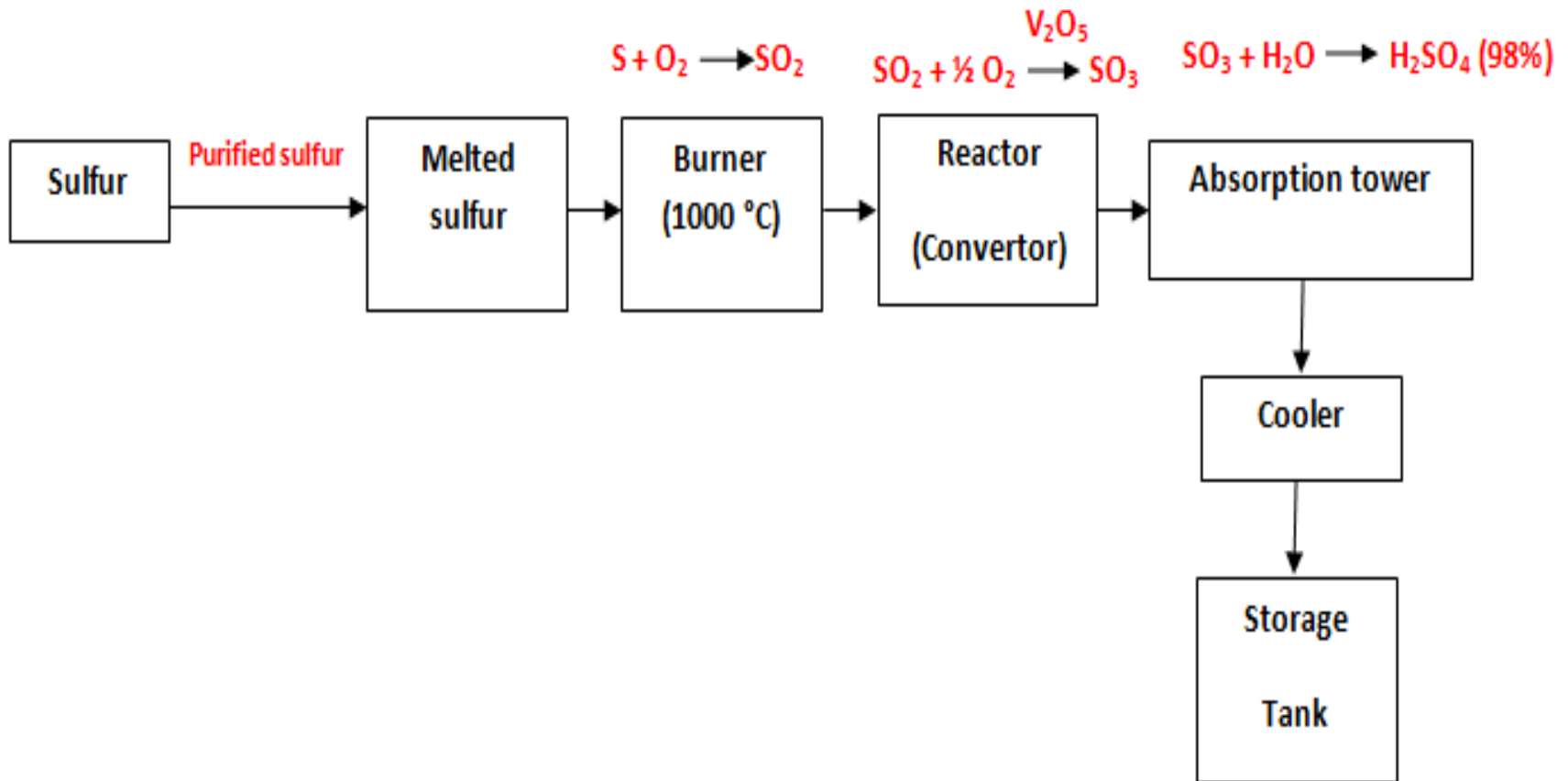
- Sulfur dioxide, cooled to ( $425^{\circ}\text{C}$ ) and passed in the state of gas into a four-stage reactor containing Vanadium oxide ( $\text{V}_2\text{O}_5$ ) as a catalyst, where it is transformed into sulfur trioxide ( $\text{SO}_3$ ).
- That gas is then directed to absorption towers where it reacts with water producing sulfuric acid at a concentration of 98.5%, then, cooled down and stored in two tanks.

## Usages:

- Sulfuric acid is used in various industries such as water treatment, batteries, and as a solvent in various industries. It is also used in producing phosphoric acid.

# Sulphuric Acid Production

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**Sulfuric acid production process**

# Environmental Hazards of $\text{H}_2\text{SO}_4$ Production

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## 1. Solid waste

- A. Spent Vanadium pentoxide ( $\text{V}_2\text{O}_5$ ) should be carefully dealt since is considered as hazardous material.
  - The spent catalyst is sealed in steel or plastic containers and stocked in old mines.
- B. Sulfur generated from filtration process( filter cake).
  - This waste will be mixed with gypsum and disposed in the gypsum disposal area.

# Environmental Hazards of $\text{H}_2\text{SO}_4$ Production

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## 2. Emissions

- $\text{SO}_2$  emissions may be originate from bad conversion of  $\text{SO}_2$  to  $\text{SO}_3$ .
- $\text{SO}_3$  emissions may be originate from bad absorption efficiency.
- $\text{H}_2\text{SO}_4$  emissions may be originate from mist formation.



# Aluminum Fluoride Production

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## Production

- Aluminum fluoride is produced from fluorosilicic acid coming as waste product from the phosphoric acid plant and Aluminum hydroxide, which is imported from abroad.
- The unit includes the follow steps:  $\text{Al}(\text{OH})_3$  handling and drying, fluorosilicic acid ( $\text{H}_2\text{SiF}_6$ ) heating, reaction, silica filtration,  $\text{AlF}_3$  crystallization, filtration, drying, and bagging.

# Aluminum Fluoride Production

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- By the completion of the reaction, the resulting mixture would be consisting of the solvent aluminum fluoride, and silica sand.
- Sand is then removed by a filtration belt, and disposed onto the belt carrying gypsum from the Phosphoric Acid Unit.
- Aluminum fluoride solution is then sent to 16 crystallizers to be separated by a filtration belt, dried through roasting in special equipment using burnt diesel, and packed according to demand in 25 kg, 50 kg, or 75 kg bags as demanded.
- **Usages:**
- Aluminum fluoride is used in extracting Aluminum from its ores by lowering the melting point in the electrical cells during the manufacturing process.

# Environmental Hazards of Aluminum fluoride Production

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- Environmental features were installed in most important areas to minimize dust emissions and recycle product (like cyclone and bag filters, and scrubbers).